

Dartmouth College

Dartmouth Digital Commons

Dartmouth Scholarship

Faculty Work

6-13-2012

The Importance of the Rotor in Hydrazone-Based Molecular Switches

Xin Su

Dartmouth College

Timo Lessing

Heinrich-Heine-Universität Düsseldorf

Ivan Aprahamian

Dartmouth College

Follow this and additional works at: <https://digitalcommons.dartmouth.edu/facoa>

 Part of the [Organic Chemistry Commons](#)

Dartmouth Digital Commons Citation

Su, Xin; Lessing, Timo; and Aprahamian, Ivan, "The Importance of the Rotor in Hydrazone-Based Molecular Switches" (2012). *Dartmouth Scholarship*. 1227.

<https://digitalcommons.dartmouth.edu/facoa/1227>

This Article is brought to you for free and open access by the Faculty Work at Dartmouth Digital Commons. It has been accepted for inclusion in Dartmouth Scholarship by an authorized administrator of Dartmouth Digital Commons. For more information, please contact dartmouthdigitalcommons@groups.dartmouth.edu.

The importance of the rotor in hydrazone-based molecular switches

Xin Su¹, Timo Lessing² and Ivan Aprahamian^{*1}

Letter

Open Access

Address:

¹Department of Chemistry, Dartmouth College, Hanover, New Hampshire 03755, United States and ²Department of Chemistry, Heinrich-Heine-Universität Düsseldorf, Universitätsstr. 1, 40225 Düsseldorf, Germany

Email:

Ivan Aprahamian* - ivan.aprahamian@dartmouth.edu

* Corresponding author

Keywords:

E/Z isomerization; hydrazone; molecular switches; pH activation; structure-property analysis

Beilstein J. Org. Chem. **2012**, *8*, 872–876.

doi:10.3762/bjoc.8.98

Received: 29 February 2012

Accepted: 09 May 2012

Published: 13 June 2012

This article is part of the Thematic Series "Molecular switches and cages".

Guest Editor: D. Trauner

© 2012 Su et al; licensee Beilstein-Institut.

License and terms: see end of document.

Abstract

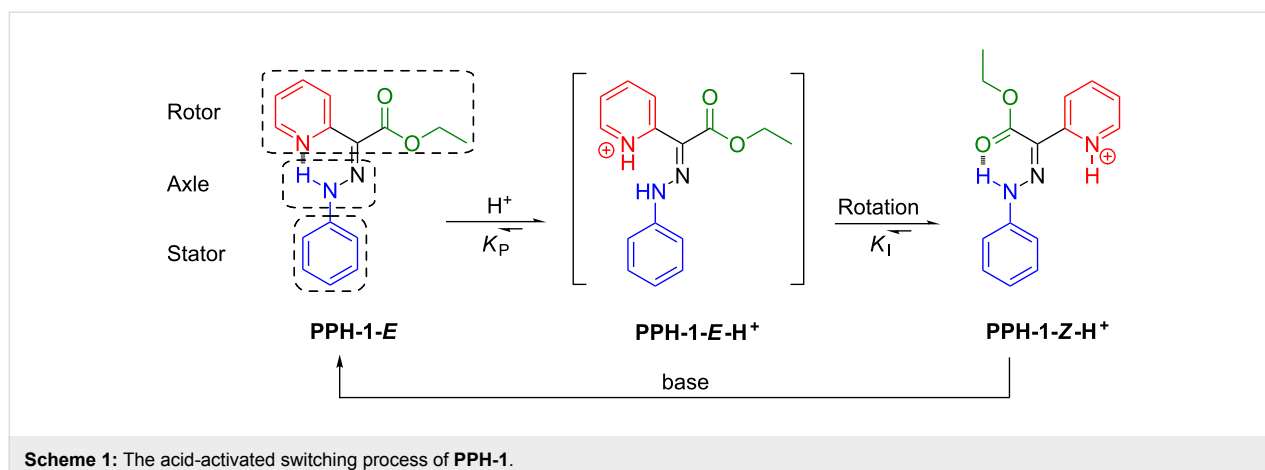
The pH-activated *E/Z* isomerization of a series of hydrazone-based systems having different functional groups as part of the rotor (R = COMe, CN, Me, H), was studied. The switching efficiency of these systems was compared to that of a hydrazone-based molecular switch (R = COOEt) whose *E/Z* isomerization is fully reversible. It was found that the nature of the R group is critical for efficient switching to occur; the R group should be a moderate H-bond acceptor in order to (i) provide enough driving force for the rotor to move upon protonation, and (ii) stabilize the obtained *Z* configuration, to achieve full conversion.

Findings

Nature is full of elegant examples of perfectly designed biological motors and machines [1] that perform delicate and precise tasks. Primitive as they may be, numerous artificial molecular machines [2-6] have been developed that strive to mimic their biological counterparts as far as function is concerned. As part of these efforts, a variety of molecular systems have been developed that can perform different types of motion (e.g., translation, rotation) in response to chemical [7-9], electrochemical [10-13], and photochemical stimuli [14-18]. One of the benefits of artificial molecular switches and machines is that their output can be controlled or fine-tuned by altering their components [19-21]. A relevant example in this context is Feringa's overcrowded alkene-based light-driven

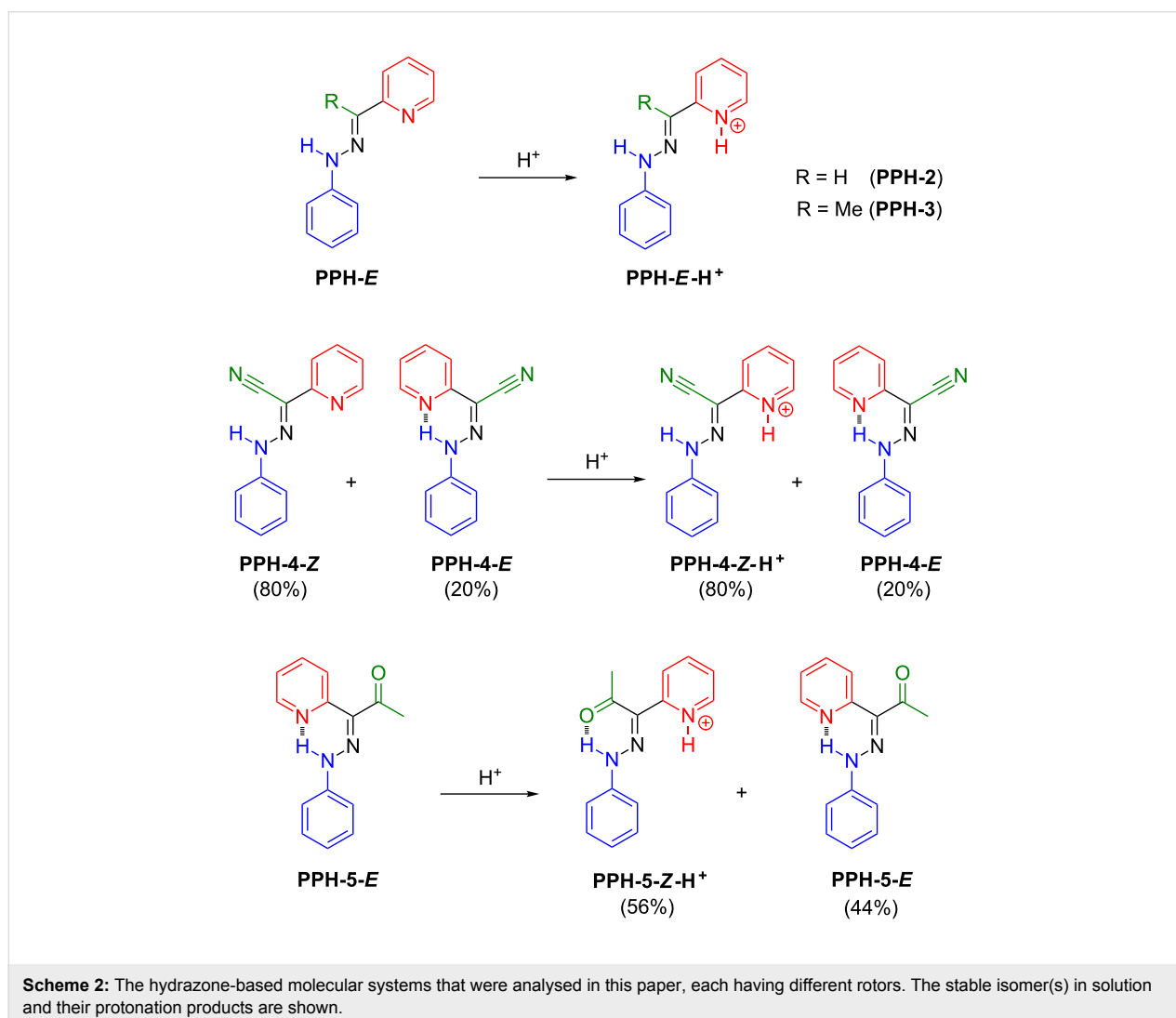
rotary switches that can be induced to rotate at different rates by replacing a naphthyl group in the upper-half of the molecule (i.e., the rotor) with a less sterically hindered benzothiophenyl group [20].

Previously, we have shown that hydrazone-based rotary switches can change their configuration (i.e., *E/Z* isomerization) as a function of pH [22-24], or upon the addition of a Lewis acid (i.e., Zn²⁺) [25]. The simplest hydrazone switch (**PPH-1**, Scheme 1) for example, exists mainly as its *E* isomer (**PPH-1-E**) in solution, as illustrated by the *E/Z* isomer ratio of 93:7 in CD₃CN. Protonation of **PPH-1-E** with acid results in an intermediate **PPH-1-E-H⁺**, which quickly isomerizes to



PPH-1-Z-H⁺, which is the more stable isomer. During this process, an *E/Z* isomerization takes place, which can be fully reversed by the addition of base to the solution.

In order to fine tune the properties of the hydrazone switches, we studied the effect of different R groups in the rotor part (Scheme 2) on the switching cycle. The target hydrazones were



synthesized either by the direct condensation of phenylhydrazine with the corresponding aldehyde (**PPH-2**) or ketone (**PPH-3**), or by Japp–Klingemann reaction (**PPH-4**, **PPH-5**) [26]. The NMR spectroscopy and mass spectrometry characterizations show results consistent with the previously reported data [26].

In certain cases it has been shown that intramolecularly H-bonded hydrazones exist predominantly as the kinetically stable *Z* isomer in solution [27–29]. We were expecting that the intramolecular H-bonds in **PPH-2** and **PPH-3** would drive them to adopt the *Z* configuration in solution as well, leading to a low-field-lying NH signal (12–16 ppm) [22–24]. However, this is not the case with **PPH-2**. The hydrazone N–H proton in **PPH-2** resonates at 8.95 ppm, which clearly shows that it is not H-bonded to the pyridyl nitrogen, indicating that the *E* configuration is the predominant isomer in solution (CD₃CN). The addition of trifluoroacetic acid (TFA) only results in a general downfield shift of the aromatic and the hydrazone N–H proton signals as a result of protonation, which reaches saturation with 3 equiv of TFA. Unlike in **PPH-1**, signals from other species are not observed during the course of protonation, suggesting that the protonation of **PPH-2** with TFA is a fast equilibrium, and that, as expected, it does not cause any isomerization. Similar to **PPH-2**, the ¹H NMR spectrum of **PPH-3** shows a signal for the hydrazone N–H proton at 8.24 ppm indicating that it too is in the *E* configuration. The protonation of **PPH-3** with TFA is a fast equilibrium as well, without any indication of rotary motion (i.e., isomerization).

On the other hand, **PPH-4**, in which R is a strong electron-withdrawing group (–CN) shows two sets of signals in the ¹H NMR spectrum (CD₃CN), indicating that two isomers, having a 4:1 ratio, coexist in solution. The major isomer shows a hydrazone N–H signal at 9.60 ppm, indicating that it is the *Z* isomer, in which an intramolecular H-bond is not present. On the other hand, the hydrazone N–H signal of the minor isomer resonates at 15.12 ppm, which is characteristic of H-bonded N–H signals, suggesting that the minor isomer is actually the *E* configuration. Such an unusual *E/Z* isomer ratio was reported before for similar systems, and it was attributed to kinetic stability of the *Z* isomer, in addition to solvent effects [27–29]. The titration of **PPH-4** with TFA only affects the major isomer (*Z*), while the minor isomer (*E*) remains intact even in the presence of 10 equiv of TFA. The changes in the ¹H NMR spectrum of **PPH-4-Z** are similar to those of **PPH-2** and **PPH-3**, except for the fact that it requires an excess of TFA (ca. 10 equiv) for the protonation to reach saturation. This observation can be attributed to the strong electron-withdrawing nature of the CN group, which drastically decreases the basicity of the pyridyl nitrogen. Furthermore, since the pyridyl nitrogen in **PPH-4-E** is

H-bonded to the hydrazone N–H, the basicity of **PPH-4-E** becomes even lower, which explains why **PPH-4-E** does not become protonated even in the presence of 10 equiv of TFA.

Structurally, **PPH-5** is the closest to **PPH-1**, that is, instead of an acyl ester group, **PPH-5** has an acetyl residue as the R group. The ¹H NMR spectrum of **PPH-5** in CD₃CN shows only one set of signals, and a sharp singlet at 14.54 ppm for the hydrazone N–H proton, indicating that it is H-bonded to the pyridyl nitrogen. Since the acetyl group is a less effective H-bond acceptor than ethyl ester, it is reasonable that **PPH-5** exists exclusively in the *E* form in solution. When TFA is added to the solution, a second set of signals arises, which grows as the amount of acid increases. The protonation of the pyridyl ring results in the downfield shift of the aromatic signals, except for proton H1, which shifts from 8.92 to 8.70 ppm as it is no longer affected by the H-bond [22–24]. Moreover, the hydrazone N–H signal shifts to a higher field (13.22 ppm) in congruence with what is observed in **PPH-1** [24]. These changes are consistent with those observed during the acid-activated switching of **PPH-1**, suggesting that **PPH-5** switches from the *E* to the *Z* configuration upon protonation. However, the switching process of **PPH-5** is relatively inefficient as there is still ca. 44% of **PPH-5-E** remaining in solution even when 30 equiv of TFA is added.

In order to rationalize the different behaviour of the structurally similar switches, **PPH-1** and **PPH-5**, a quantitative evaluation of the thermodynamic process is necessary. Taking a look at the acid-activated switching process of **PPH-1** (Scheme 1), we can formulate the following equations for the acid-induced *E/Z* isomerization:

$$K_P = \frac{[\text{PPH-1-E-H}^+]}{[\text{PPH-1-E-H}][\text{H}^+]} \quad (1)$$

$$K_I = \frac{[\text{PPH-1-Z-H}^+]}{[\text{PPH-1-E-H}^+]} \quad (2)$$

$$K_S = K_P \cdot K_I = \frac{[\text{PPH-1-Z-H}^+]}{[\text{PPH-1-E-H}][\text{H}^+]} \quad (3)$$

where K_P is the equilibrium constant of the protonation step, K_I is the equilibrium constant for the rotation process, and K_S is the overall equilibrium constant for the switching reaction. The $\text{p}K_a$ of **PPH-1** is actually $\log_{10}K_P$, so K_S also equals $10^{[\text{p}K_a(\text{PPH-1})]}$. K_I . From the above equations, it becomes clear that K_S can be used as an index to evaluate the feasibility of the

switching process in hydrazone-based switches; the larger the K_S value, the easier the switching process. In the case of **PPH-1** versus **PPH-5**, since the acetyl group is a stronger electron-withdrawing group than the ester group, the basicity (pK_a) of the pyridyl group in **PPH-1** will be higher than in **PPH-5**. Moreover, the ester group is a better H-bond acceptor than the acetyl group, which means that the protonated *Z* configuration of **PPH-1** is more stable than that of **PPH-5**, resulting in a larger K_I for **PPH-1**. Thus, it can be qualitatively deduced that **PPH-1** has a larger K_S than **PPH-5**, suggesting that **PPH-1** is a more ideal system to be used as a molecular switch. This analysis is clearly in line with the acid switching experiments that show that **PPH-1** can be fully switched, whereas **PPH-5** cannot.

Conclusion

In summary, we have synthesized four hydrazone-based systems having different R groups as part of the rotor section. The role of the R group was assessed *vis-à-vis* the switching of the system, and it was found that for the switch to operate effectively it is crucial that (1) the R group be able to offer a second H-bond-accepting site in order to provide enough driving force for the rotor to move; and (2) the R group be a moderate H-bond acceptor, otherwise the isomer generated will not be stable enough to enable full conversion (isomerization).

Supporting Information

Supporting Information File 1

Experimental section and acid titration of the hydrazone compounds.

[<http://www.beilstein-journals.org/bjoc/content/supplementary/1860-5397-8-98-S1.pdf>]

Acknowledgements

This research was supported by Dartmouth College and the Burke Research Initiation Award. Timo Lessing wishes to thank the German Academic Exchange Service (DAAD) for a RISE worldwide research internship.

References

- Schilwa, M., Ed. *Molecular Motors*; Wiley-VCH: Weinheim, Germany, 2003.
- Kelly, T. R.; De Silva, H.; Silva, R. A. *Nature* **1999**, *401*, 150–152. doi:10.1038/43639
- Leigh, D. A.; Wong, J. K. Y.; Dehez, F.; Zerbetto, F. *Nature* **2003**, *424*, 174–179. doi:10.1038/nature01758
- Fletcher, S. P.; Dumur, F.; Pollard, M. M.; Feringa, B. L. *Science* **2005**, *310*, 80–82. doi:10.1126/science.1117090
- Kelly, T. R.; Cai, X.; Damkaci, F.; Panicker, S. B.; Tu, B.; Bushell, S. M.; Cornella, I.; Piggott, M. J.; Salives, R.; Cavero, M.; Zhao, Y.; Jasmin, S. *J. Am. Chem. Soc.* **2007**, *129*, 376–386. doi:10.1021/ja066044a
- Feringa, B. L. *J. Org. Chem.* **2007**, *72*, 6635–6652. doi:10.1021/jo070394d
- Badjić, J. D.; Balzani, V.; Credi, A.; Silvi, S.; Stoddart, J. F. *Science* **2004**, *303*, 1845–1849. doi:10.1126/science.1094791
- Berná, J.; Leigh, D. A.; Lubomska, M.; Mendoza, S. M.; Pérez, E. M.; Rudolf, P.; Teobaldi, G.; Zerbetto, F. *Nat. Mater.* **2005**, *4*, 704–710. doi:10.1038/nmat1455
- Bonnet, S.; Collin, J.-P.; Koizumi, M.; Mobian, P.; Sauvage, J.-P. *Adv. Mater.* **2006**, *18*, 1239–1250. doi:10.1002/adma.200502394
- Wang, W.; Kaifer, A. E. *Angew. Chem., Int. Ed.* **2006**, *45*, 7042–7046. doi:10.1002/anie.200602220
- Nijhuis, C. A.; Ravoo, B. J.; Huskens, J.; Reinhoudt, D. N. *Coord. Chem. Rev.* **2007**, *251*, 1761–1780. doi:10.1016/j.ccr.2007.02.001
- Juluri, B. K.; Kumar, A. S.; Liu, Y.; Ye, T.; Yang, Y.-W.; Flood, A. H.; Fang, L.; Stoddart, J. F.; Weiss, P. S.; Huang, T. J. *ACS Nano* **2009**, *3*, 291–300. doi:10.1021/nn8002373
- Parimal, K.; Witlicki, E. H.; Flood, A. H. *Angew. Chem., Int. Ed.* **2010**, *49*, 4628–4632. doi:10.1002/anie.201001003
- Balzani, V.; Credi, A.; Venturi, M. *Pure Appl. Chem.* **2003**, *75*, 541–547. doi:10.1351/pac200375050541
- Kottas, G. S.; Clarke, L. I.; Horinek, D.; Michl, J. *Chem. Rev.* **2005**, *105*, 1281–1376. doi:10.1021/cr0300993
- Eelkema, R.; Pollard, M. M.; Vicario, J.; Katsonis, N.; Ramon, B. S.; Bastiaansen, C. W. M.; Broer, D. J.; Feringa, B. L. *Nature* **2006**, *440*, 163. doi:10.1038/440163a
- Muraoka, T.; Kinbara, K.; Aida, T. *Nature* **2006**, *440*, 512–515. doi:10.1038/nature04635
- Russev, M.-M.; Hecht, S. *Adv. Mater.* **2010**, *22*, 3348–3360. doi:10.1002/adma.200904102
- Kulago, A. A.; Mes, E. M.; Klok, M.; Meetsma, A.; Brouwer, A. M.; Feringa, B. L. *J. Org. Chem.* **2010**, *75*, 666–679. doi:10.1021/jo902207x
- Landaluce, T. F.; London, G.; Pollard, M. M.; Rudolf, P.; Feringa, B. L. *J. Org. Chem.* **2010**, *75*, 5323–5325. doi:10.1021/jo1006976
- Lubbe, A. S.; Ruangsapichat, N.; Caroli, G.; Feringa, B. L. *J. Org. Chem.* **2011**, *76*, 8599–8610. doi:10.1021/jo201583z
- Landge, S. M.; Aprahamian, I. *J. Am. Chem. Soc.* **2009**, *131*, 18269–18271. doi:10.1021/ja909149z
- Su, X.; Aprahamian, I. *Org. Lett.* **2011**, *13*, 30–33. doi:10.1021/ol102422h
- Landge, S. M.; Tkatchouk, E.; Benítez, D.; Lanfranchi, D. A.; Elhabiri, M.; Goddard, W. A., III; Aprahamian, I. *J. Am. Chem. Soc.* **2011**, *133*, 9812–9823. doi:10.1021/ja200699v
- Su, X.; Robbins, T. F.; Aprahamian, I. *Angew. Chem., Int. Ed.* **2011**, *50*, 1841–1844. doi:10.1002/anie.201006982
- Yang, Y.; Su, X.; Carroll, C. N.; Aprahamian, I. *Chem. Sci.* **2012**, *3*, 610–613. doi:10.1039/c1sc00658d
- Karabatsos, G. J.; Shapiro, B. L.; Vane, F. M.; Fleming, J. S.; Ratka, J. S. *J. Am. Chem. Soc.* **1963**, *85*, 2784–2788. doi:10.1021/ja00901a025
- Mitchell, A. D.; Nonhebel, D. C. *Tetrahedron Lett.* **1975**, *16*, 3859–3862. doi:10.1016/S0040-4039(00)91297-8
- Butler, R. N.; Johnson, S. M. *J. Chem. Soc., Perkin Trans. 1* **1984**, 2109–2116. doi:10.1039/P19840002109

License and Terms

This is an Open Access article under the terms of the Creative Commons Attribution License (<http://creativecommons.org/licenses/by/2.0>), which permits unrestricted use, distribution, and reproduction in any medium, provided the original work is properly cited.

The license is subject to the *Beilstein Journal of Organic Chemistry* terms and conditions: (<http://www.beilstein-journals.org/bjoc>)

The definitive version of this article is the electronic one which can be found at:
[doi:10.3762/bjoc.8.98](https://doi.org/10.3762/bjoc.8.98)